POTENTIOMETRIC STUDY OF RHENIUM(V) COMPLEX FORMATION WITH AZATHIOPRINE AND CEFTRIAXONE

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Received: 3 March 2017; Accepted: 17 September 2017

Abstract

The behavior of complex formation of Re(V) metal ion with azathioprine (AZ) and ceftriaxone (CE) as medicinal drugs has been investigated potentiometrically in 0.3 M hydrochloric acid. The stability constant of the prepared complexes has been determined and calculated at various temperatures. The obtained results showed that Re(V)-AZ system showed the formation 1:2, 1:3 and 1:4 complexes, while Re(V)-CE showed only 1:1 and 1:2 complexes. The complex formation suggests a successive displacement of the chloride molecules from the coordination sphere of the central ion by the AZ ligand. In addition, the effect of temperature on the complexes formation of Re(V) with AZ and CE was investigated. The entropy and enthalpy changes showed a favorable and exothermic process, respectively. The kinetic parameters of the formation complex process were calculated and discussed.

Keywords: rhenium(V), drugs, potentiometry, complex formation, stability constant

Introduction

The stability of complexes with medicinal drugs plays a major role in their biological and chemical activities. This is due to that, the medicinal drugs have various functional groups, which can bind to metal ions present in the human body [1, 2]. The azathioprine (AZ) is an immunosuppressive drug used in organ transplantation and autoimmune diseases [3 – 5] while the ceftriaxone (CE) is effective against a wide variety of Gram-positive and Gram-negative bacteria [6 – 9]. Very few works investigated the stability constant and thermodynamic parameters (ΔG, ΔH and ΔS) of Re-Az and Re-CE complexes at different temperatures. The complex formation behavior of AZ with various metal ions has been studied by the potentiometric titration method [10]. AZ is found to form 2:1
complexes with Co(II), Cu(II) and Ni(II), in the stability order of Cu(II) > Ni(II) > Co(II). The effect of various
divalent cations on the encapsulation efficiency of gellan gum and to probe the underlying mechanisms responsible
for drug loading efficiency. Spherical beads containing AZ were prepared from deacetylated gellan gum by
ionotropic gelation method [11]. The anti-rheumatic studies on AZ have also involved its Cu(II) complexes as the
possible active form and the stoichiometry and the stability constants of AZ metal complexes, even with that of
Cu(II), have been determined [12].

CE is evaluated as a corrosion inhibitor for carbon steel alloys in the presence of 0.5 M hydrochloric acid (HCl) at
different concentrations, i.e., 10–50 ppm and different temperatures [13]. Studies of the mechanism, kinetics and
thermodynamic constants of VO(II) and Cu(II) complexes with CE in the pH range of 2 – 5 at temperatures of 303,
313, 323 and 333 K, using reversed phase HPLC and ion exchange HPLC were reported [14]. CE degradation in
aqueous solution was studied at 310 and 353 K in acidic, basic and neutral conditions. Stability constants of Re(V)
metal complexes with selected medicinal drugs adenosine, isoniazid and metformin hydrochloride have been
determined using a pH metric titration technique in aqueous HCl at different temperatures and an ionic strength of
0.2 M [15]. The process of complex formation of Re(V) ion with 2-mercapto-pyridine was investigated
potentiometrically by evaluating formation the equilibrium and stability constants of the metal complex at the
temperature range of 273-338 K [16].

In the present study, the formation equilibrium investigations were extended to the stoichiometry and stability
constants of complex formation of AZ and CE (Scheme 1) with Re(V) metal ion in aqueous HCl at different
temperatures. Moreover, the kinetic parameters of the complex formation were determined and discussed.

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**Scheme 1.** The chemical structures of the studied drugs

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**Materials and Methods**

**Materials**

Potassium perrhenate (KReO₄) and iodine were purchased from Aldrich and Fluka, respectively. KReO₄ was
converted to K₂ReOCl₅ precursor according to the method reported elsewhere [17]. The concentration of metal ion
solution was determined spectrophotometrically at wavelength (λ) of 480 nm. A pure AZ and CE drugs were
purchased from the local market (El-Nasr Pharm. Chem. Co., Egypt), potassium iodide and HCl (36.4%) was
purchased from BDH.

**Titration method**

Potentiometric titrations were performed using a R3003 voltage comparator. A platinum plate was used as an
indicator electrode. The redox system consisting of AZ and CE was created by oxidizing a small portion of the
initial drugs in 0.3 M HCl with 0.1 N iodine solution (0.5 mL). The stepwise complexes of Re(V) with AZ and CE
were studied using the Bjerrum method [18, 19]. For this purpose, the redox system consisting of the drugs
(0.04192 mol/L) and its oxidized form titrated with a K₂ReOCl₅ solution (0.0799 mol/L) in 0.3 M HCl. The
equilibrium ligand concentration ([L]) and the formation function (\(\bar{n}\)) at each titration point was calculated using the equations (1) and (2) [20, 21], respectively:

\[
\log [L] = \frac{E_I - E_i}{1.987 \times 10^{-4} T} + \log C_L^1 + \frac{1}{2} \log \frac{V_I}{V_T} \\
\bar{n} = \frac{C_L^1 - [L]}{C_M}
\] (1)  (2)

where, \(E_I\) and \(E_i\) are the initial equilibrium potentials of the system in the absence of metal ion and at the current titration point, \(C_L^1\) is the drug initial concentration of, \(V_I\) and \(V_T\) are the initial and total volumes, respectively, \(T\) is the temperature in K and \(C_M\) is the concentration of metal ion at each titration point.

The stepwise stability constants (log \(K_i\)) of the complexes were calculated at half-integer values of the degree of formation. Thus, a large equilibrium constant indicates a highly stable complex. The overall stability constant (\(\beta_i\)) defined as the following equation 3 [22]:

\[
\log \beta_i = \log K_1 + \log K_2 + \log K_3 \ldots \ldots + \log K_i
\] (3)

The log \(K_i\) of the complexes was calculated at 288, 298, 308 and 318 K.

The complex formation thermodynamic characteristics such as: enthalpy (\(\Delta H\)), Gibb’s free energy (\(\Delta G\)) and entropy (\(\Delta S\)) are a fundamental important properties of understanding of the various factors such as electronic and steric effects, solute-solvent interactions that may influence on the complexes [23]. The thermodynamic parameters of the metal ion complexes at 288, 298, 308, and 318 K were studied. The complex degree of formation was obtained according to the equations (4) and (5) [24, 25]:

\[
\Delta G = \Delta H - T \Delta S
\] (4)

\[
\log K_i = \frac{-\Delta H}{2.303 R} + \frac{1}{T} \times \frac{\Delta S}{2.303 R}
\] (5)

where, \(R\) is the ideal gas constant (8.314 J/K mol). \(\Delta H\) and \(\Delta S\) were obtained from the intercept and slope of the plot of \(\log K_i\) against \(1/T\).

**Results and Discussion**

**Potentiometric titration**

The metal ion in solution excites as combined with ligands or chelating groups instead of isolated form, giving rise to complexions or coordination compounds. The obtained titration data were analyzed and then stability constant values were calculated [26]. The experimental data show that the equilibrium potential of the redox system increases in the process of potentiometric titration as the metal ion volume increases. During the titration, no precipitates were formed indicating that there is no tendency to form hydroxide complexes [27]. Intense coloration was observed, which indicated the complexes formation. The formation curves were plotted between \(\bar{n}\) and -log [L]. The coordination process in the systems Re(V)-AZ and Re(V)-CE in 0.3 M HCl at various temperatures were characterized by the formation curves as displayed in Figure 1.

Re(V) exists in solutions and solids as octahedral species [28, 29]. In this study, the obtained complexes of Re(V) with AZ and CE have a coordination number of six [30, 31]. AZ is considered as neutral and monodentate ligand and the coordination bonds occur between the lone pair of sulfur atom and Re(V) [32]. The Re(V)-AZ system has a maximum value of \(\bar{n} = 4\), this indicates the formation 1:2, 1:3 and 1:4 complexes. While about 1:1 complex has not formed. The complex formation suggests a successive displacement of the chloride molecules from the coordination sphere of the central ion by the AZ and CE ligands [33]. The maximum value of \(\bar{n}\) in the Re(V)-CE system was 2, this shows the formation of 1:1 and 1:2 complexes. On the other hand, CE loses one sodium atom and becomes anionic. In case of 1:1 complex, Re(V) unit is coordinated by a mono anionic O,O-bidentate unit of CE ligand, one
CE and three halogens, with Re(V) center in a distorted octahedral environment. The halide ligands are in cis positions to each other [34, 35]. The proposed mechanism of the Re(V)-CE complex formation process 1:1 and 1:2 ratios according to the following reaction (Scheme 2). The estimated log K values of Re(V) complexes with AZ and CE at different temperatures are listed in Table 1.

![Formation curves of Re(V) complexes with AZ and CE in 0.3 M HCl at 318 K](image)

Figure 1. Formation curves of Re(V) complexes with AZ and CE in 0.3 M HCl at 318 K

![Scheme 2. The chelation between the [ReOCl3]2- and CE](image)
Table 1. Stepwise stability constants of Re(V) complexes in 0.3 M HCl

<table>
<thead>
<tr>
<th>T (K)</th>
<th>log $K_{2AZ}$</th>
<th>log $K_{3AZ}$</th>
<th>log $K_{4AZ}$</th>
<th>log $\beta_{iAZ}$</th>
<th>log $K_{1CE}$</th>
<th>log $K_{2CE}$</th>
<th>log $\beta_{iCE}$</th>
</tr>
</thead>
<tbody>
<tr>
<td>288</td>
<td>5.93</td>
<td>5.30</td>
<td>1.82</td>
<td>13.05</td>
<td>6.66</td>
<td>4.95</td>
<td>11.61</td>
</tr>
<tr>
<td>298</td>
<td>5.84</td>
<td>5.13</td>
<td>1.77</td>
<td>12.72</td>
<td>6.34</td>
<td>4.67</td>
<td>11.01</td>
</tr>
<tr>
<td>308</td>
<td>5.32</td>
<td>4.66</td>
<td>1.72</td>
<td>11.63</td>
<td>6.11</td>
<td>4.33</td>
<td>10.44</td>
</tr>
<tr>
<td>318</td>
<td>4.91</td>
<td>4.31</td>
<td>1.66</td>
<td>10.84</td>
<td>5.35</td>
<td>4.10</td>
<td>9.45</td>
</tr>
</tbody>
</table>

It is observed from Table 1 that, the temperature increase leads to decreasing in log $\beta$ values for both AZ and CE. It could be seen from data in all the cases that the differences between $K_1$, $K_2$, $K_3$ and $K_4$ values are found to be high which indicate the formation of the stepwise complex. The value of the ratio of log $K_i$/log $K_j$ or log $K_j$/log $K_3$ is positive in all the cases. This implies that there is no steric hindrance to the addition of a secondary ligand molecule [36]. This is due to the statistically effective and statistically coordination of a second molecule is difficult when compared to the first due to the availability of less number of coordinating sites on the metal ion for the second ligand [37].

**Complexes thermodynamic**

The chemical species in the reaction undergo a change in their concentrations, at the same time. The Gibb's free energy is a function of the concentrations of reactants and products. Entropy is a measure of energy that is unavailable for useful, chemical work. The entropy of an individual species is always positive and tends to be larger for gases than for solids, and for more complex molecules than for simple molecules. The thermodynamic parameters values of Re(V)-AZ and Re(V)-CE complexes have been calculated from the temperature dependent data given in Table 1.

The thermodynamic data were plotted as the values of equilibrium constants (log $K_i$) versus the reciprocal of temperature (1/T) as shown in Figures 2 and 3 for AZ and CE complexes, respectively. All the thermodynamic parameters of the stepwise stability constants of the complexes are given in Table 2.

![Figure 2](https://example.com/figure2.png)

**Figure 2.** The relations between the values of equilibrium constants (log $K_i$) versus the reciprocal of temperature (1/T) of Re(V)-AZ complexes: log $K_2$ corresponds to [ReO(AZ)$_2$Cl$_3$]; log $K_3$ corresponds to [ReO(AZ)$_3$Cl$_2$]$^+$ and log $K_4$ corresponds to [ReO(AZ)$_4$Cl]$^{5+}$ species
Figure 3. The relations between the values of equilibrium constants (log $K_i$) versus the reciprocal of temperature $(1/T)$ of Re(V)-CE complexes: log $K_1$ corresponds to [ReO(CE)Cl$_3$] and log $K_2$ corresponds to [ReO(CE)$_2$Cl] species.

Table 2. Thermodynamic parameters of Re(V) complexes with AZ and CE in 0.3 M HCl

<table>
<thead>
<tr>
<th>Species</th>
<th>$\Delta G$ (kJ/mol)</th>
<th>$\Delta H$ (kJ/mol)</th>
<th>$\Delta S$ (J/K mol)</th>
</tr>
</thead>
<tbody>
<tr>
<td>[ReO(AZ)$_2$Cl$_3$]</td>
<td>32.32</td>
<td>62.36</td>
<td>-100.78</td>
</tr>
<tr>
<td>[ReO(AZ)$_3$Cl$_3$]$^+$</td>
<td>28.58</td>
<td>60.07</td>
<td>-105.66</td>
</tr>
<tr>
<td>[ReO(AZ)$_3$Cl]$_2$$^{2+}$</td>
<td>10.08</td>
<td>92.76</td>
<td>2.71</td>
</tr>
<tr>
<td>[ReO(CE)Cl$_3$]</td>
<td>35.99</td>
<td>72.40</td>
<td>-12.18</td>
</tr>
<tr>
<td>[ReO(CE)$_2$Cl]</td>
<td>26.51</td>
<td>50.70</td>
<td>-81.60</td>
</tr>
</tbody>
</table>

The stepwise formation constants of complexes of the type $ML_n$ (where $M$ is the metal, $L$ is the ligand and $n$ is the number of ligands) decrease in magnitude as $n$ increases. This may be partly explained in terms of the entropy factor. In the case of the formation of octahedral complexes of Re(V)-AZ, the reaction between $K_2$[ReOCl$_5$] and AZ involves three successive reactions as the following equations (6-8).

$$K_2[ReOCl_5] + 2AZ \rightleftharpoons [ReO(AZ)$_2$Cl$_3$] + 2KCl \quad (6)$$

$$[ReO(AZ)$_2$Cl$_3$] + AZ \rightleftharpoons [ReO(AZ)$_3$Cl$_2$]Cl \quad (7)$$

$$[ReO(AZ)$_3$Cl$_2$]Cl + AZ \rightleftharpoons [ReO(AZ)$_4$Cl]_2 \quad (8)$$

In equation (6), the two ligands ($n = 2$) can go into two cites out of five of $K_2$[ReOCl$_5$]. For equation (8), $n = 4$ and the fourth ligand can go into one of two sites of $K_2$[ReOCl$_5$]. This means that there is more randomness in the first step than the fourth one, so $\Delta S$ is more positive (2.71), $\Delta H$ is more negative (-92.76) and log $K_2 >$ log $K_4$ [38]. In addition, there is a dramatic increase in $\Delta S$ for adding two compared to adding four monodentate AZ ligands (-100.78 to 2.71 J/K mol) which changed from negative (unfavorable) to positive (favorable). Using the equilibrium constant for the complex formation of Re(V)-CE where the one and two bidentates ligand replace the monodentates chloride ions, it has been found the $\Delta S$ decrease (-12.18 to -81.60). The negative change in $\Delta S$ indicates a highly solvated metal complex and indicated that the formation of these complexes was entropy favored [39, 40]. $\Delta H$
change for the complexes suggests that all the complexation reactions releasing heat have a negative value (exothermic), favorable at a lower temperature. The metal-ligand binding process is enthalpy driven and metal-ligand bonds are strong [41]. The sign of ΔG indicates the direction in which a reaction moves to reach its equilibrium position. The complexation reactions are thermodynamically favorable when its enthalpy decreases and its entropy increase. Substituting the inequalities ΔH < 0 and ΔS > 0 into the previous equation (4) shows that complexation reactions are thermodynamically favorable and spontaneous when ΔG is negative [42]. Similar trend was obtained for Re(V) with another ligand [38].

**Distribution diagrams**

Mole fractions are commonly used to calculate the concentrations of the individual complexes based on the formation constants. Mole fractions of a form of the complex compressed as the ratio of the concentration of the complex to the total concentration of the metal ion. The distribution diagrams were drawn in the titration where the Re(V) to ligands, mole ratio was 1:1, 1:2, 1:3 and 1:4 at 318 K (Figure 4). It was obtained using Microsoft Excel calculations [43] and the concentration of total metal ions (0.0799 mol/L) was set as 100%.

In Figure 4, the species [ReO(AZ)Cl] for the system Re(V)-AZ reached a maximum of 2.9%. The second species [ReO(AZ)Cl] and [ReO(AZ)Cl] reached a maximum of 85.3% and 96.6%, respectively. For Re(V)-CE system, it reached a maximum of 67.5% and 99.89% for [ReO(CE)Cl] and [ReO(CE)Cl] species, respectively. The analysis of the temperature-dependent distribution function shows that the yield of the complexes decreases with increasing the temperature. The analysis of the distribution curves provides the possibility to elucidate the predominance region for some complex depending on temperature and concentration. These results were used to develop optimal methods for the preparation of Re(V) metal ion complexes with AZ and CE in 0.3 M HCl.

![Distribution curves of Re(V) complexes with AZ and CE at 318 K](image)

**Conclusion**

The potentiometric study investigated the behavior of complex formation of Re(V) metal ion with azathioprine (AZ) and ceftriaxone (CE). AZ is a neutral and monodentate ligand and the coordination bonds occur between the lone pair of sulfur atom with Re(V). Re(V)-AZ system has a maximum value of $\bar{n}$ of 4, this showed the formation 1:2, 1:3 and 1:4 complexes. The complex formation suggests a successive displacement of the chloride molecules from the coordination sphere of the central ion by the AZ ligand. The maximum value of $\bar{n}$ in the system of Re(V)-CE was 2, this showed the formation 1:1 and 1:2 complexes. The CE is losing one sodium atom and becomes anionic.
ligand. The increasing amount coordinated molecules AZ and CE and the temperature the values log $K_i$ decreases. There is a dramatic increase $\Delta S$ for adding two compared to adding four monodentate AZ ligands (-100.78 to 2.71 J/K mol) which change from negative (unfavorable) to positive (favorable). The enthalpy change ($\Delta H$) for the complexes suggests that all the complexation reactions releasing heat have a negative $\Delta H$ (exothermic), favorable at low temperatures. $\Delta G$ indicates the direction in which a reaction moves to reach its equilibrium position. The above mentioned parameters and results will be useful in understanding the biological behavior of these complexes in the biological applications.

Acknowledgement
The authors would like to thank Chemistry department, Faculty of Science, Al-Azhar University, Assiut, Egypt for providing the facilities of this work.

References


