



SYNTHESIS, CHARACTERISATION AND EFFECT OF TEMPERATURE ON CORROSION INHIBITION BY THIOSEMICARBAZONE DERIVATIVES AND ITS TIN(IV) COMPLEXES

(Sintesis, Pencirian dan Kesan Suhu Ke atas Perencatan Kakisan oleh Ligan Terbitan Tiosemikarbazon dan Kompleks Timah(IV))

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Abstract

Thiosemicarbazone derivatives and its tin(IV) complexes used in this study were 2-acetylpyridine 4-ethyl-3-thiosemicarbazone (HAcETSc), 2-acetylpyridine 4-methyl-3-thiosemicarbazone (HAcMTSc), 2-acetylpyridine 4-ethyl-3-thiosemicarbazone dichlorophenyltin (Sn(HAcETSc)PhenCl₂), and 2-acetylpyridine 4-methyl-3-thiosemicarbazone dichlorophenyltin (Sn(HAcMTSc)PhenCl₂). All these title compounds were characterised using an elemental analyser, Fourier-transform infrared-attenuated total reflectance spectroscopy (FTIR-ATR), ultraviolet-visible spectroscopy (UV-Vis), and nuclear magnetic resonance (NMR). The efficiency of the synthesised compounds as a corrosion inhibitor of mild steel in 1.0 M HCl was investigated using weight loss technique at different concentrations, 1, 2, and 3 mM, and at a temperature range of 30–60 °C. The thiosemicarbazone ligand and its tin(IV) complexes affirmed the hypothesis where the inhibitor efficiency tends to increase as the inhibitor concentration increases, indicating their potential use as a corrosion inhibitor for mild steel. Moreover, inhibitor efficiency decreases when temperature increases.

Keywords: thiosemicarbazone, corrosion inhibitor, temperature, mild steel

Abstrak

Terbitan tiosemikarbazon dan kompleks timah(IV) yang digunakan dalam kajian ini adalah 2-asetilpiridina 4-etil-3-tiosemikarbazon (HAcETSc), 2-asetilpiridina 4-metil-3-tiosemikarbazon (HAcMTSc), 2-asetilpiridina 4-etil-3-tiosemikarbazon dikloridastanum (Sn(HAcETSc)PhenCl₂) and 2-asetilpiridina 4-metil-3-tiosemikarbazon dikloridastanum (Sn(HAcMTSc)PhenCl₂). Kesemua sebatian telah dicirikan menggunakan analisis unsur, spektroskopi inframerah transformasi Fourier-pantulan keseluruhan dikeselikan (FTIR-ATR), spektroskopi ultralembayung-sinar nampak (UV-Vis) dan resonans magnetik nuklear (NMR). Kecekapan sebatian yang telah disintesis dikaji sebagai perencat kakisan terhadap keluli lembut dalam 1.0 M HCl dengan kepekatan berbeza, 1, 2, dan 3 mM, serta suhu antara 30-60 °C menggunakan kaedah kesusutan berat. Ligan tiosemikarbazon dan kompleks timah(IV) membuktikan hipotesis di mana kecekapan perencat cenderung untuk meningkat apabila kepekatan perencat meningkat, menunjukkan ia berpotensi untuk digunakan sebagai perencat kakisan terhadap keluli lembut. Tambahan lagi, kecekapan perencat menurun apabila suhu meningkat.

Kata kunci: tiosemikarbazon, perencatan kakisan, suhu, keluli lembut

Introduction

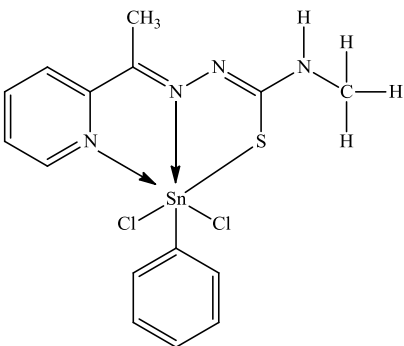
Corrosion inhibitor is one of the methods used to prevent corrosion [1]. Inhibitors are compounds used in low concentration and applied in an aggressive environment to protect or slow down corrosion [2, 3]. Finding new corrosion inhibitors, which have high efficiency, are easy-to-make, and cheap, have attracted many researchers' attention [4]. In the industrial sector, building surface corrosion occurs from common activities such as acid pickling, acid descaling, and industrial acid cleaning, where an acid solution, such as hydrochloric acid, is used to remove undesirable scale or rust. This activity causes metal dissolution due to the attack of the aggressive media. A corrosion inhibitor containing heteroatoms of nitrogen (N), sulphur (S), and oxygen (O), as well as an aromatic group in its structure, is considered effective and has the potential as a corrosion inhibitor [5, 6]. Through adsorption process, the inhibitor forms a protective layer and offers a barrier from corrosive attack on the metal surface [7, 8].

Thiosemicarbazone (TSC) derivatives and tin complexes have been reported to be potential corrosion inhibitors due to the presence of nitrogen and sulphur atoms, which could increase inhibition efficiency [9]. Furthermore, tin complexes with thiosemicarbazone derivatives have received a great attention as a corrosion inhibitor, since organotin is recognised for its biocidal behaviour, friction and wear reducing properties, as well as corrosion-inhibiting characteristics [10,11]. Therefore, this study aims to synthesise and characterise new compounds for their expected anticorrosion activity. The synthesised ligands and complexes were tested as potential corrosion inhibitors in 1.0 M HCl. The compounds under investigation are listed in Table 1.

Table 1. Compounds under investigation

Inhibitor	Abbreviation	Structure	Molar Mass (g/mol)
2-acetylpyridine 4-ethyl-3-thiosemicarbazone	HAcETSc		222
2-acetylpyridine 4-methyl-3-thiosemicarbazone	HAcMTSc		208
2-acetylpyridine 4-ethyl-3-thiosemicarbazone dichlorophenyltin	Sn(HAcETSc)PhenCl ₂		486

Table 1 (cont'd). Compounds under investigation

Inhibitor	Abbreviation	Structure	Molar Mass (g/mol)
2-acetylpyridine 4-methyl-3-thiosemicarbazone dichlorophenyltin	Sn(HAcMTSc)PhenCl ₂		472

Materials and Methods

Material and measurement

All chemicals used were commercially purchased without further purification. The melting point was measured using a melting point apparatus (model SMP10 Stuart) and was taken in an open capillary tube. The elemental analysis of synthesised compounds was performed using CHNS/O Model Fision EA 1180 and Thermo Finnigan Flash EA 1112 Series. Meanwhile, the infrared spectra were recorded on Fourier-transform infrared-attenuated total reflectance (FTIR-ATR) Perkin Elmer spectrophotometer. The electronic absorption spectra were recorded on a PG Instrument T80/T80+ spectrophotometer in the 200–600 nm region using dimethyl sulfoxide (DMSO) as a solvent. ¹H and ¹³C NMR spectra were measured on a Bruker spectrophotometer using deuterated dimethyl sulfoxide (DMSO-d₆) as a solvent and tetramethylsilane as an internal reference.

Synthesis of ligands

An ethanoic solution of 2-acetylpyridine (1 mmol) and ethanoic solution of 4-ethyl-3-thiosemicarbazide or 4-methyl-3-thiosemicarbazide (1 mmol) were mixed together, and then a few drops of acetic acid were added to the mixture. The mixture was refluxed for 2 h with a constant stirring rate of 400 rpm at a temperature between 60–70 °C. The product was cooled after 2 h and kept at room temperature to obtain a crystalline product. The crystalline product formed was filtered, washed with cold ethanol, and dried over silica gel.

Synthesis of tin(IV) complexes

The synthesised ligand (1 mmol) was dissolved in ethanol and the metal salt, phenyltin trichloride (1 mmol), was dissolved in distilled water. The metal salt was added dropwise to the ethanoic solution ligand with constant stirring at 400 rpm. The reaction mixture was continuously stirred for 4 hours. The precipitate was then filtered and washed with cold ethanol. The precipitate formed was left to dry over silica gel. The weight of the precipitate was recorded.

Corrosion inhibition study - weight loss technique

The aggressive solution, 1 M hydrochloric acid (HCl), 37% of analytical grade was prepared *via* dilution with distilled water. The inhibitor with concentrations of 1, 2, and 3 mM were also prepared by diluting the stock solution, 1 M HCl. The determination of weight loss on mild steel was tested for 24 hours at 30 °C. The experiment was conducted in triplicate and the average mass was calculated. The experiment was repeated at different temperatures of 40, 50, and 60 °C.

Results and Discussion

Physical measurement

The physical characteristics of the synthesised ligands and complexes, namely yield percentage, colour, melting point, molar conductivity, and elemental percentage are shown in Table 2. The complexes have a higher melting point compared to the ligands. The molar conductivities for both complexes are in the range of 4.0–12.5 $\text{Scm}^2\cdot\text{mol}^{-1}$, which indicates non-electrolyte behaviour [12]. The elemental analysis of calculated and experimental data is in good agreement.

Table 2. Physical characteristics of synthesised ligands and complexes

Compound	Percentage Yield (%)	Colour	Melting Points (°C)	Molar Conductivity ($\text{Scm}^2\cdot\text{mol}^{-1}$)	Elemental Analysis (Experimental) (%)			
					C	H	N	S
HAcETSc	88.06	Light yellow	127-130	-	54.05 (53.34)	6.31 (7.41)	25.23 (24.50)	14.42 (14.67)
HAcMTSc	71.61	Light yellow	176-177	-	51.56 (51.92)	5.88 (5.77)	28.06 (26.92)	14.55 (15.38)
Sn(HAcETSc)PhenCl ₂	63.37	Yellow	232-235	7.5	39.51 (39.94)	3.70 (3.81)	11.52 (11.91)	6.58 (6.34)
Sn(HAcMTSc)PhenCl ₂	56.14	Yellow	242-245	5.2	38.14 (37.59)	3.39 (3.39)	11.86 (11.83)	6.78 (6.28)

Infrared spectra

The tentative assignments to the significant IR spectral bands of ligands and complexes are presented in Table 3 and Table 4, respectively. The infrared spectra for compounds HAcETSc and Sn(HAcETSc)PhenCl₂ are shown in Figure 1, whereas the infrared spectra for compounds HAcMTSc and Sn(HAcMTSc)PhenCl₂ are shown in Figure 2. Thiosemicarbazone ligand contains thiomide, a C=S functional group, which is known to exist as a thione-thiol tautomerism. Nonetheless, the absence of S-H stretching band of 2500–2600 cm^{-1} indicates that the ligand is in a thione form [13]. Several significant changes in spectral complexes and ligands suggest the coordination of TSC ligands to metal ion through azomethine nitrogen and thiolate sulphur. In the spectra of ligands, the stretching bands at 3210 and 3240 cm^{-1} for compounds HAcETSc and HAcMTSc, respectively, correspond to the $\nu(\text{N-H})$ group that disappeared, and a new stretching band appeared in the spectra of complexes, which is attributed to $\nu(\text{C=N})$ stretching band due to thioenolisation of C=S [14].

Moreover, changes can be observed through the shifting of some bands. The shifting of $\nu(\text{C=N}_{\text{py}})$ stretching band to higher frequency from the spectra of ligand indicates the involvement of pyridine azomethine nitrogen in coordination with metal ion. Besides, $\nu(\text{C=N})$ observed at 1527 and 1537 cm^{-1} in both ligands shifted to 1504 and 1557 cm^{-1} , respectively, in the spectra of complexes, indicating that the coordination occurred through azomethine nitrogen.

The stretching bands due to $\nu(\text{C=S})$ at 1084 cm^{-1} [compound HAcETSc] and 1074 cm^{-1} [compound HAcMTSc] disappeared in the spectra of complexes and new stretching bands, C-S appeared at 1061 cm^{-1} [compound Sn(HAcETSc)PhenCl₂] and 1047 cm^{-1} [compound Sn(HAcMTSc)PhenCl₂], indicating that the coordination occurred through thiolate sulphur after enolisation of C=S functional group and followed by deprotonation on sulphur [15]. Nevertheless, the new stretching bands at 310–360 cm^{-1} , 258–327 cm^{-1} , and 330–354 cm^{-1} were assigned to $\nu(\text{Sn-N})$, $\nu(\text{Sn-N}_{\text{py}})$, and $\nu(\text{Sn-S})$, respectively, and the stretching band at 229–270 cm^{-1} range was assigned to $\nu(\text{Sn-Cl})$ [16]. The existence of these new stretching bands indicates the coordination of ligands to metal ion.

Table 3. IR spectral data of the ligand and complexes (cm^{-1})

Compound	$\nu(\text{N-H})$	$\nu(\text{C=N}_{\text{py}})$	$\nu(\text{C=N})$	$\nu(\text{N-N})$	$\nu(\text{C-N})$	$\nu(\text{C=S})$	$\nu(\text{C-S})$
HAcETSc	3349, 3210	1580	1527	1297	1221	1084	-
HAcMTSc	3288, 3240	1578	1537	1295	1230	1074	-
$\text{Sn}(\text{HAcETSc})\text{PhenCl}_2$	3372	1601	1504, 1587	1283	1247	-	1061
$\text{Sn}(\text{HAcMTSc})\text{PhenCl}_2$	3297	1601	1557, 1587	1290	1256	-	1047

Table 4. Far-IR spectral data of the complexes (cm^{-1})

Compound	$\nu(\text{Sn-N})$	$\nu(\text{Sn-S})$	$\nu(\text{Sn-N}_{\text{py}})$	$\nu(\text{Sn-Cl})$
$\text{Sn}(\text{HAcETSc})\text{PhenCl}_2$	331	354	289	267
$\text{Sn}(\text{HAcMTSc})\text{PhenCl}_2$	328	357	289	265

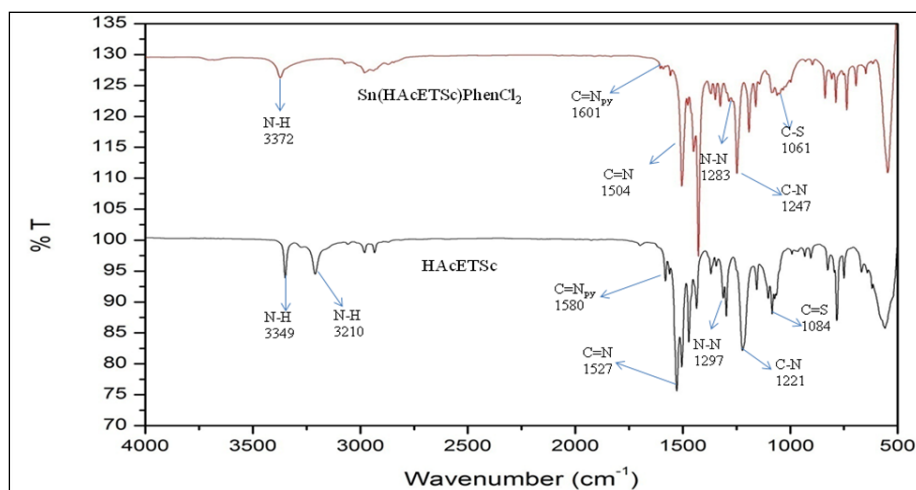


Figure 1. Infrared spectra for compound HAcETSc and $\text{Sn}(\text{HAcETSc})\text{PhenCl}_2$

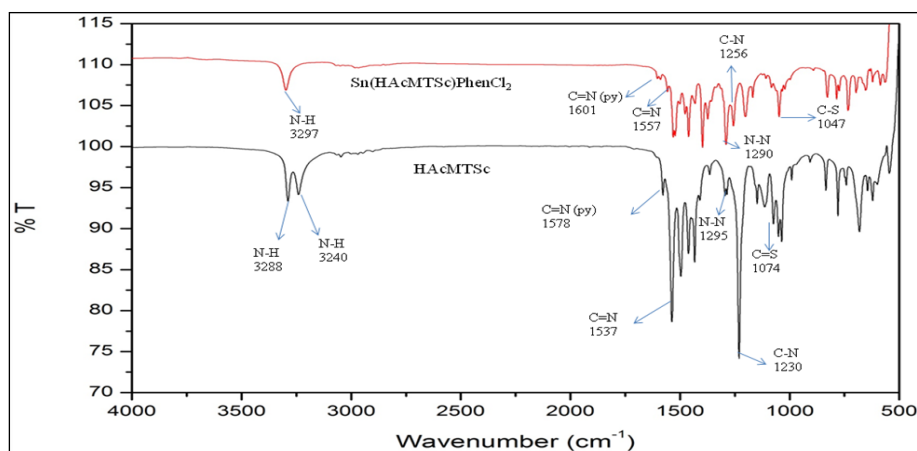


Figure 2. Infrared spectra for compound HAcMTSc and $\text{Sn}(\text{HAcMTSc})\text{PhenCl}_2$

Electronic spectra

The UV-Visible (UV-Vis) absorption spectra for compounds HAcETSc and Sn(HAcETSc)PhenCl₂ are shown in Figure 3, whereas the UV-Vis absorption spectra for compounds HAcMTSc and Sn(HAcMTSc)PhenCl₂ are shown in Figure 4. For both ligands, there is only one absorption peak at 320 nm and assigned to n→π* transition. This band corresponds to the C=S and C=N groups. The absorption peak attributed to π→π* transition might not be observed in the ligand spectrum as n→π* transition in C=S and C=N has stronger absorption and might cover π→π* transition [17]. Upon complexation, two absorption peaks were observed at 315 and 400 nm [compound Sn(HAcETSc)PhenCl₂], as well as at 305 and 405 nm [compound Sn(HAcMTSc)PhenCl₂], which were appointed to n→π* transition and the ligand to metal charge transfer (LMCT), respectively. Moreover, both complexes underwent a hypsochromic shift. The spectra shifting shows the donation of electron lone pair from N or S to the metal ion [18].

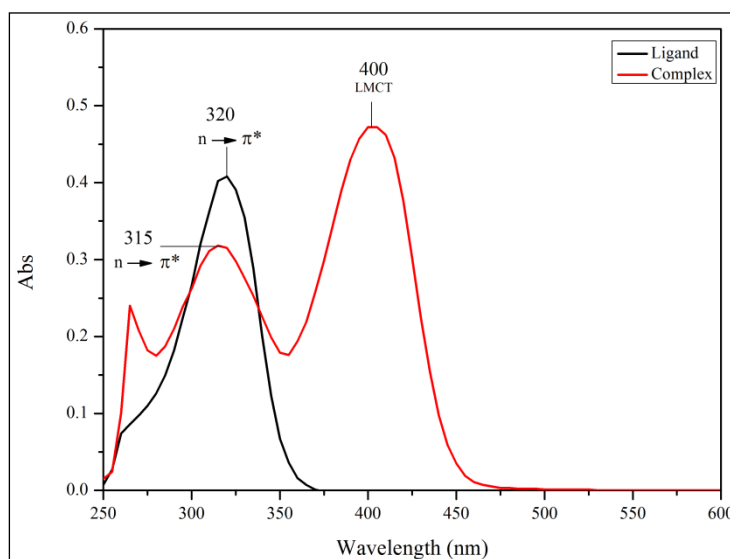


Figure 3. UV-Vis absorption spectra for compound HAcETSc and Sn(HAcETSc)PhenCl₂

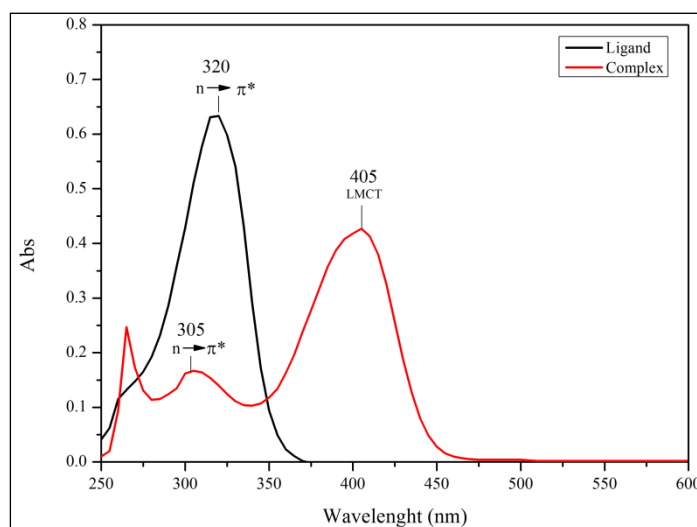


Figure 4. UV-Vis absorption spectra for compound HAcMTSc and Sn(HAcMTSc)PhenCl₂

¹H and ¹³C NMR spectra

¹H and ¹³C NMR spectral data of the ligands and complexes are tabulated in Table 5 and 6, respectively. For ligands HAcETSc and HAcMTSc, N(1)H signal appeared as a triplet at 8.68 and 8.62 ppm, respectively. Meanwhile, N(2)H signal appeared as a singlet at 10.25 ppm for both ligands. Upon complexation, the peak of N(1)H shifted downfield with a small shifting value and N(2)H signal disappeared due to deprotonation of the ligands when coordinated to the metal ion. Furthermore, CH₃-C=N signal of complexes showed about 0.4 ppm shift downfield as compared to the ligand, which revealed the involvement of an azomethine nitrogen in coordination to the metal ion. Multiplets signal were found for aromatic ring proton. The aromatic ring proton signals shifted downfield upon complexation with metal [19].

In ¹³C NMR, the signal of C=S and C=N groups shifted to the upfield region in the complexes as compared to the ligands, which proved that the coordination occurred through N=C-S⁻ and nitrogen atom of pyridine group to the metal ion. Moreover, the coordination to metal ion occurred *via* an azomethine nitrogen was proven when the CH₃-C=N signal of the complexes shifted to downfield region, in comparison with the ligands. There were extra signals present, which are represented as a phenyl ring [20].

Table 5. ¹H NMR spectral data (ppm) of the ligands and complexes

Compound	Chemical shift, δ (ppm)					
	N(1)H	N(2)H	-CH ₃	CH ₃ -C=N	-CH ₂	Aromatic
HAcETSc	8.68	10.25	1.16	2.38	3.64	7.36 - 8.58
HAcMTSc	8.62	10.25	3.05	2.38	-	7.37 - 8.59
Sn(HAcETSc)PhenCl ₂	8.70	-	1.22	2.78	3.55	7.52 - 8.70
Sn(HAcMTSc)PhenCl ₂	8.68	-	3.02	2.79	-	7.51 - 8.68

Table 6. ¹³C NMR spectral data (ppm) of the ligands complexes

Compound	Chemical shift, δ (ppm)					
	C=S	C=N	-CH ₃	CH ₃ C=N	-CH ₂	Phenyl Ring
HAcETSc	178.10	155.17	14.93	12.59	39.01	121.26 - 148.93
HAcMTSc	179.18	155.22	31.63	12.56	-	121.24 - 148.93
Sn(HAcETSc)PhenCl ₂	168.76	135.77	15.05	14.46	37.80	142.24 - 144.87, 125.91 - 130.71
Sn(HAcMTSc)PhenCl ₂	169.43	135.76	29.63	15.10	-	142.52 - 144.86, 125.96 - 130.73

Corrosion inhibition study

The inhibitor efficiency of mild steel in 1 M HCl at different temperatures and different concentrations was investigated by weight loss method, in the presence and absence of inhibitors. The values of inhibitor efficiency are given in Table 7. The results showed that the inhibition efficiency increased when the concentration of inhibitor increased. Apart from that, the title compounds showed a decreasing efficiency at higher temperature (60 °C). The results were different at certain temperatures for each compound. For instance, at concentration 3 mM, compound HAcETSc had better efficiency at 50 °C while other compounds worked better at 40 °C.

Table 7. Inhibitor efficiency at different concentrations in 1M HCl at different temperatures

Concentration	Compounds	Inhibitor Efficiency (%)			
		30 °C	40 °C	50 °C	60 °C
3 mM	HAcETSc	84.74	91.09	96.08	84.80
	HAcMTSc	84.32	95.91	83.80	80.48
	Sn(HAcETSc)PhenCl ₂	86.44	89.88	81.61	73.32
	Sn(HAcMTSc)PhenCl ₂	94.78	95.98	95.46	87.33
2 mM	HAcETSc	84.32	88.94	85.27	84.10
	HAcMTSc	78.67	78.00	81.06	76.10
	Sn(HAcETSc)PhenCl ₂	79.52	87.32	71.84	40.95
	Sn(HAcMTSc)PhenCl ₂	85.03	94.57	92.30	53.58
1 mM	HAcETSc	81.92	77.48	82.93	67.61
	HAcMTSc	42.51	66.03	72.32	61.56
	Sn(HAcETSc)PhenCl ₂	54.52	80.14	57.58	34.76
	Sn(HAcMTSc)PhenCl ₂	81.53	84.38	91.41	45.22

At 30 °C, the highest percentage of inhibitor efficiency is as follows: Sn(HAcMTSc)PhenCl₂ > Sn(HAcETSc)PhenCl₂ > HAcETSc > HAcMTSc. The efficiency for compound Sn(HAcMTSc)PhenCl₂ at 30 °C was 94.78%; the value increased slightly at 40 °C and started to decrease until 60 °C with the efficiency of 87.33%. Similarly, for compounds HAcMTSc and Sn(HAcETSc)PhenCl₂, its efficiency increased at 40 °C but became less efficient with increasing temperature. For compound HAcETSc, the efficiency kept increasing until the temperature was 50 °C and decreased at 60 °C. Even though the efficiency is low at higher temperature, it is still able to protect mild steel in the acid solution from corroding as compared to the blank of 1 M HCl [21].

Generally, the effects of a compound to provide better protection on mild steel surface are due to the presence of electron pairs, which are sulphur and nitrogen atoms, π -electrons on the aromatic rings, molecular size, and the formation of metallic complexes [22, 23]. Hence, it can be recommended that compound Sn(HAcMTSc)-PhenCl₂ is efficient in preventing mild steel dissolution as compared to other studied compounds

Conclusion

The thiosemicarbazone derivative ligands derived from 2-acetylpyridine and tin(IV) complexes have been successfully synthesised through condensation process. The ligands and complexes were characterised by using CHNS, FTIR, UV-Vis, and NMR. From the characterisation of the complexes, it has been determined that ligand acts as tridentate to metal ions, which coordinates through azomethine nitrogen and thiolate sulphur. The outcome for the corrosion study showed good inhibitive properties for mild steel corrosion and the inhibitor efficiency increased as the inhibitor concentration increased. For temperature effect, the inhibitor efficiency decreased at high temperature. This work suggests that the corrosion inhibition study should undergo further study using electrochemical impedance spectroscopy (EIS) and linear polarisation resistance (LPR) to confirm the corrosion efficiency based on the study done using weight loss technique.

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